

Class Xii Chemistry Practical Salt Analysis

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This is done by conducting a series of tests in a systematic manner and using the observations to confirm the absence or presence of specific cations and anions. Salt analysis is an integral part of the CBSE class 12 chemistry practical examinations and is a topic that several students struggle with.

Salt Analysis: Step-by-Step Guide with Tricks & Shortcuts

$N_2 = 1$ (for Mohr's salt), $V_2 = 10\text{ml}$, $M_2 = 1/20\text{M}$. $M_1 = [1 \cdot (1/20) \cdot 10] / [5 \cdot 8.8] = 1/88\text{M} = 0.01\text{M}$. Strength = $M \times \text{Molar Mass} = 158 \cdot (1/88) = 1.79\text{g/L}$. RESULT- (ON RULED SIDE)- The Molarity of $\text{KMnO}_4 = 0.01\text{M}$. And the strength of $\text{KMnO}_4 = 1.79\text{g/L}$. You can also get Class XII Practicals on Biology, Physics, Chemistry, and Physical Education.

CBSE Class XII Practicals - Titration of Mohr's Salt

In a test tube, take 0.1 g of salt, add a pinch of manganese dioxide and 3-4 drops of conc. Acid with sulfuric acid. Heat the mixture of reactions. Greenish yellow chlorine gas that is detected by its strong odor and bleaching action. 8. Br⁻-(Bromide anion) Take 0.1 g of salt in a test tube with a pinch of MnO_2 .

Systematic Analysis of Anions - Chemistry Practicals Class 12

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