

# Chapter 11 Stoichiometry

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## Chapter 11 Stoichiometry

Stoichiometry is the tool for answering these questions. Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass.

## Chapter 11: Stoichiometry

Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass. Recall that the law states that matter is neither created nor destroyed in a chemical reaction.

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In Section 11.3, for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate

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volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water).

## **Chapter 11.4: Stoichiometry - Chemistry LibreTexts**

15.2 CHAPTER 11: STOICHIOMETRY MOLE TO MOLE RATIO When nitrogen and hydrogen gas are heated under the correct conditions, ammonia gas ( $\text{NH}_3$ ) is formed.

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Chapter 12 . Stoichiometry Notes Packet Big Picture Ideas: The identity of the reactants helps scientists to predict the products in a chemical reaction. Quantitative relationships exist with all chemical reactions that allow scientists to predict amounts of products formed, reactants consumed, and percent yield based on theoretical maximum.

## **CHAPTER 11: STOICHIOMETRY**

stoichiometry The study of quantitative relationships between the amounts of reactants used and the amounts of products formed. It is based on the law of conservation of mass (the mass of the reactants equals the mass of the products).

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Solutions Manual Chemistry: Matter and Change • Chapter 11 209 Stoichiometry Stoichiometry CHAPTER 11 SOLUTIONS MANUAL Section 11.1 Defining Stoichiometry pages 368–372 Practice Problems pages 371–372 1. Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that the law of conservation of mass is

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Chapter 11: Stoichiometry. Chemistry. Section 11.1: Defining Stoichiometry. Particle and Mole Relationships: \_\_\_\_\_ is the study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction.

### **Chapter 11: Stoichiometry**

This is a standard stoichiometry problem of the type presented in Section 11.4, except this problem asks for the volume of one of the reactants (O<sub>2</sub>) rather than its mass. We proceed exactly as in Section 11.4, using the strategy

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Combustion of liquid propane also releases more energy than natural gas. Ratiospage 964 Math Handbook Rhonda Peacher Photography <http://glencoe.com> 376 Chapter 11 Stoichiometry EXAMPLE Problem 11.3 Mole-to-Mass Stoichiometry Determine the mass of sodium chloride (NaCl), commonly called table salt, produced when 1.25 mol of chlorine gas (Cl<sub>2</sub>) reacts vigorously with excess sodium. 1 Analyze the Problem You are given the moles of the reactant, Cl<sub>2</sub>, and must determine the mass of the product, NaCl.

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